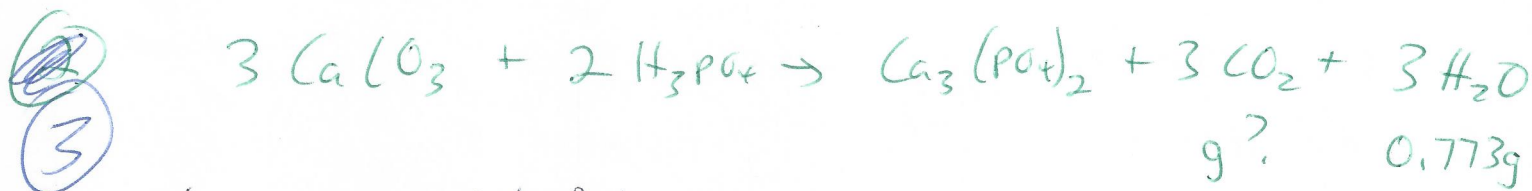


$$(5.0\text{g CaC}_2) \left(\frac{1\text{ mol CaC}_2}{64\text{ g CaC}_2} \right) = 0.078\text{ mol CaC}_2$$

$$(0.078\text{ mol CaC}_2) \left(\frac{1\text{ mol C}_2\text{H}_2}{1\text{ mol CaC}_2} \right) = 0.078\text{ mol C}_2\text{H}_2$$

$$(0.078\text{ mol C}_2\text{H}_2) \left(\frac{26\text{ g C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} \right) = 2.03\text{ g C}_2\text{H}_2$$

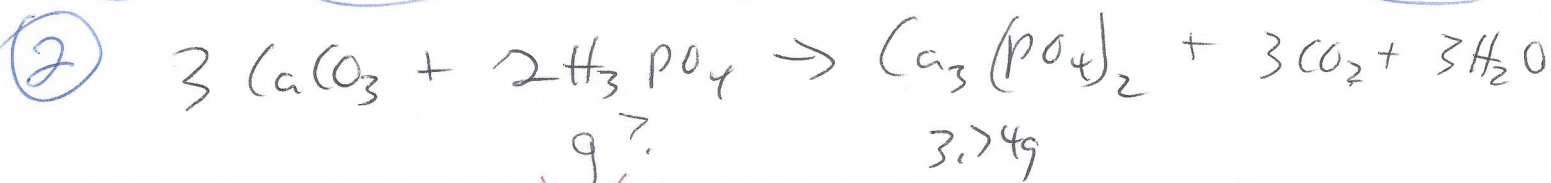


$$(0.773\text{g H}_2\text{O}) \left(\frac{1\text{ mol H}_2\text{O}}{18\text{ g H}_2\text{O}} \right) = 0.043\text{ mol H}_2\text{O}$$

$$(0.043\text{ mol H}_2\text{O}) \left(\frac{3\text{ mol CO}_2}{3\text{ mol H}_2\text{O}} \right) = 0.043\text{ mol CO}_2$$

$$(0.043\text{ mol CO}_2) \left(\frac{44\text{ g CO}_2}{1\text{ mol CO}_2} \right) = 1.9\text{ g CO}_2$$

out of order
happens
when you
get
interrupted!



$$(3.74\text{g Ca}_3(\text{PO}_4)_2) \left(\frac{1\text{ mol Ca}_3(\text{PO}_4)_2}{310\text{ g Ca}_3(\text{PO}_4)_2} \right) = 0.0121\text{ mol Ca}_3(\text{PO}_4)_2$$

$$(0.0121\text{ mol Ca}_3(\text{PO}_4)_2) \left(\frac{2\text{ mol H}_3\text{PO}_4}{1\text{ mol Ca}_3(\text{PO}_4)_2} \right) = 0.0242\text{ mol H}_3\text{PO}_4$$

$$(0.0242\text{ mol H}_3\text{PO}_4) \left(\frac{98\text{ g H}_3\text{PO}_4}{1\text{ mol H}_3\text{PO}_4} \right) = 2.37\text{ g H}_3\text{PO}_4$$

